

---

This is an electronic reprint of the original article.  
This reprint may differ from the original in pagination and typographic detail.

Kanninen, Petri; Nguyen Dang, Luong; Le, Hoang Sinh; Anoshkin, Ilya V.; Tsapenko, Alexey; Seppälä, Jukka; Nasibulin, Albert G.; Kallio, Tanja

**Transparent and flexible high-performance supercapacitors based on single-walled carbon nanotube films**

*Published in:*  
Nanotechnology

*DOI:*  
[10.1088/0957-4484/27/23/235403](https://doi.org/10.1088/0957-4484/27/23/235403)

Published: 28/04/2016

*Document Version*  
Peer-reviewed accepted author manuscript, also known as Final accepted manuscript or Post-print

*Please cite the original version:*  
Kanninen, P., Nguyen Dang, L., Le, H. S., Anoshkin, I. V., Tsapenko, A., Seppälä, J., Nasibulin, A. G., & Kallio, T. (2016). Transparent and flexible high-performance supercapacitors based on single-walled carbon nanotube films. *Nanotechnology*, 27(23), 1-6. Article 235403. <https://doi.org/10.1088/0957-4484/27/23/235403>

**Transparent and flexible high-performance supercapacitors based on single-walled carbon nanotube films**

Petri Kanninen<sup>1</sup>, Nguyen Dang Luong<sup>2</sup>, Le Hoang Sinh<sup>2,3</sup>, Ilya V. Anoshkin<sup>4,5</sup>, Alexey Tsapenko<sup>6</sup>, Jukka Seppälä<sup>2</sup>, Albert G. Nasibulin<sup>4,6,7</sup>, Tanja Kallio<sup>1</sup>

<sup>1</sup> Research Group of Electrochemical Energy Conversion and Storage, Department of Chemistry, School of Chemical Technology, Aalto University, P.O. Box 16100, FI-00076 Aalto, Finland.

<sup>2</sup> Laboratory of Polymer Technology, Department of Biotechnology and Chemical Technology, School of Chemical Technology, Aalto University, P.O. Box 16100, FI-00076 Aalto, Finland.

<sup>3</sup> Institute of Research and Development, Duy Tan University, R.809, K7/25 Quang Trung, Da Nang, Vietnam.

<sup>4</sup> Department of Applied Physics, School of Science, Aalto University, P.O. Box 15100, FI-00076 Aalto, Finland.

<sup>5</sup> Department of Radio Science and Engineering, School of Electrical Engineering, Aalto University, P.O. Box 13000, FI-00076 Aalto, Finland.

<sup>6</sup> Skolkovo Institute of Science and Technology, 100 Novaya str., Skolkovo, Moscow Region, 143025, Russia.

<sup>7</sup> Saint-Petersburg State Polytechnical University, Department of Material Science, Polytechnicheskaya 29, 195251, Saint-Petersburg, Russia.

---

Email: [a.nasibulin@skoltech.ru](mailto:a.nasibulin@skoltech.ru) and [tanja.kallio@aalto.fi](mailto:tanja.kallio@aalto.fi)

## **Abstract**

Transparent and flexible energy storage devices have gathered great interest due to their suitability for display, sensor and photovoltaic applications. In this paper, we report the application of aerosol synthesized and dry deposited single-walled carbon nanotube (SWCNT) thin films as electrodes for electrochemical double-layer capacitor (EDLC). SWCNT films exhibit extremely large specific capacitance ( $178 \text{ F g}^{-1}$  or  $552 \text{ } \mu\text{F cm}^{-2}$ ), high optical transparency (92%) and stability for 10000 charge/discharge cycles. A transparent and flexible EDLC prototype is constructed with a polyethylene casing and a gel electrolyte.

## **Keywords**

Energy storage; single-walled carbon nanotube; supercapacitor; transparent; flexible

## **1. Introduction**

Electrochemical double-layer capacitors (EDLC) are supercapacitors based on high surface area carbon, where the charge is stored in the electrochemical double-layer between the electrolyte and the carbon electrode. They can store several times more charge per volume or mass than traditional electrolytic capacitors, their charge/discharge times are fast and they are very stable due to lack of chemical reactions in the cell [1]. Recently, interest in making transparent and flexible EDLCs has risen due to several applications in displays, touch sensors, photovoltaics, OLED, *etc.*, which hold promise to change the way we use electronics. Carbon nanotubes and other carbon materials can be used to fabricate electrodes for these novel transparent and flexible EDLC (Table 1) [2-15]. However, to achieve high optical transparency one should use ultrathin films, which limits the conductivity of the electrodes and sets a requirement for a very high mass specific capacitance to achieve adequate performance. In this respect, single-walled carbon nanotubes (SWCNT) are an ideal material for transparent and flexible EDLCs due to their excellent electric conductivity and specific surface area.

SWCNTs have many unique properties, which are advantageous for a wide variety of applications including EDLCs. They have exceptionally high Young's modulus of elasticity and tensile strength and are the strongest known material [16,17]. The porosity and specific surface area of SWCNT films are very large and they possess high transparency and flexibility. The SWCNTs can withstand extremely high currents (up to  $10^9 \text{ A cm}^{-2}$ ) making them an ideal replacement for copper and aluminum in fast integrated charge/discharge circuits. Due to these properties SWCNT films are thought to be excellent replacement for traditional transparent electrodes, like indium-tin oxide (ITO), which suffers from several serious drawbacks including limited flexibility and chemical resistance, rare raw material and high refraction and haze. SWCNT networks show improved performance and reduced fabrication cost compared to ITO and organic materials that have been studied as low-cost alternatives [18].

Current techniques for SWCNT thin film preparation are restricted in various ways. In substrate chemical vapor deposition (CVD) the high temperature of SWCNT growth limits the choice of substrates. Solution-based method is a time and resource consuming process with multiple steps including purification and dispersion of SWCNT, which will produce films with defected SWCNT due to ultrasonic treatment. Therefore, aerosol CVD process was recently combined with SWCNT dry deposition to solve these problems [19,20]. The SWCNTs grown by ferrocene decomposition in CO atmosphere are filtered on low adhesive force filters at the outlet. From there, the thin film can be transferred effectively to any substrate, *e.g.* flexible and transparent separator materials for EDLCs. SWCNT produced with the aerosol method fulfil all the requirements for low-cost, flexible and transparent electronics that can be used in EDLCs [21-23].

In this work, dry deposited thin films of aerosol CVD synthesized SWCNT are used as the electrodes in EDLCs. Remarkably high specific capacitance and stability are reached with the SWCNT electrodes in test cells and they are used to construct a transparent and flexible EDLC prototype.

**Table 1.** Literature review of flexible and transparent supercapacitors based on carbon nanomaterials. Specific capacitances ( $C_{sp}$ ) are normalized to one electrode according to Eq. 2 and 3 from the values reported in the references.

Material	Cell type	Transparency @ 550 nm (%)	$C_{sp}$ $F g^{-1}$	$C_{sp}$ $\mu F cm^{-2}$	Reference
SWCNT	test cell	92	178	552	This work
CNF-RGO	Prototype	30		1730	[2]
CNT	Prototype	75	29.2		[3]
CNT	3-electrode	N/A	60		[4]
graphene	prototype	57	7.6	5.8	[5]
graphene	prototype	84		8.4	[6]
graphene	prototype	64		24.8	[7]
graphene	prototype	73	17.3	5.33	[8]
graphene	3-electrode	70	135		[9]
MWCNT	prototype	75	100		[10]
MWCNT	prototype	70	6.38		[11]
nanocups	prototype	71		818	[12]
SWCNT	3-electrode	54	12.9		[13]
SWCNT	prototype	60	90	584	[14]
SWCNT	prototype	58	136.8		[15]

## 2. Experimental

SWCNTs were synthesized by an aerosol (floating catalyst) CVD method based on ferrocene vapor decomposition in a CO atmosphere described elsewhere [24,25]. Briefly, the catalyst precursor was vaporized by passing ambient temperature CO through a cartridge filled with ferrocene powder. The flow containing ferrocene vapor was then introduced into the high-temperature zone. In order to obtain stable growth of SWCNTs, a controlled amount of carbon dioxide was added together with the carbon source [26]. SWCNTs were collected downstream of the reactor by filtering the flow through nitrocellulose or silver membrane filters. It is worth noting that the SWCNTs produced by this method are of a high quality, which was confirmed with transmission electron microscope (TEM) examinations and optical investigations. Thus, they are immediately suitable for many applications without any additional treatment. Nevertheless, each SWCNT contains a catalyst particle encapsulated inside the tube and covered by graphitic carbon and cannot be washed out by HCl without opening the caps of the SWCNTs. The SWCNT film on different substrates was formed by a dry press transfer technique [19]. The film mass was weighed with a Mettler Toledo XP2 microbalance.

Structural and morphological characterizations of SWCNTs were probed with TEM (JEOL JEM-2200FS operated at 80 kV) and scanning electron microscopy (SEM, Zeiss sigma VP). Raman spectra were collected with a LabRAM HR UV-NIR instrument using back-scattered light from a 633 nm laser focused through a 50× objective lens. Absorbance measurements were carried out with Perkin Elmer Lambda 950 UV-Vis-NIR spectrophotometer.

The EDLCs were fabricated from SWCNT thin films as follows: a circular piece of Celgard 2300 (thickness 25  $\mu\text{m}$ , diameter 1.5 cm) separator was cut and 1x1 cm piece of the thin film was transferred by simply pressing the separator against the film tightly. Thin films were transferred on both sides of the separator one after another. The SWCNT/separator was then dipped into 20 wt%

HNO<sub>3</sub> for 15 minutes to oxidize the SWCNTs and increase the hydrophilicity of the thin film [27]. Finally, the SWCNT/separator was washed with deionized water and placed in 1 M H<sub>2</sub>SO<sub>4</sub> electrolyte overnight to saturate.

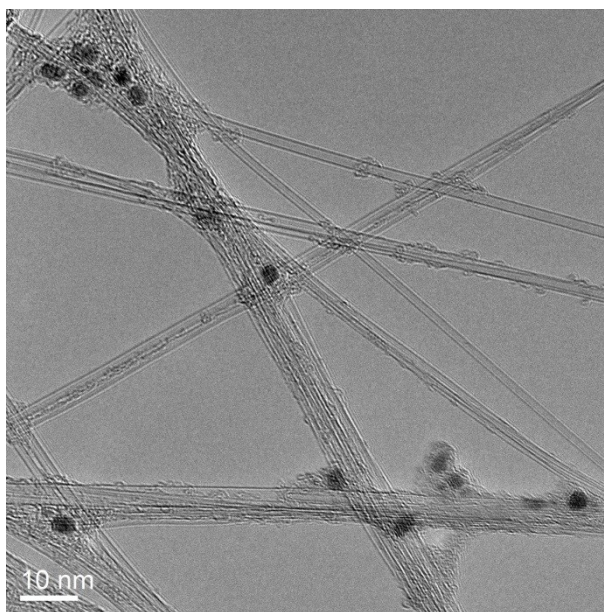
The electrochemical measurements were made in a Hohen Corp test cell and 2 drops of electrolyte were placed in the cell along with the soaked SWCNT/separator to ensure proper wetting. The assembled EDLC was let to stabilize for 1 h before the start of the measurements and it was subsequently cycled between 0 and 0.6 V for 25 times before any presented measurements were made. Cyclic voltammograms were recorded at scan speeds between 10 and 100 mV s<sup>-1</sup>. Constant current charging and discharging was performed at 0.53 A g<sup>-1</sup> from 0 to 0.6 V and backward. The stability of the EDLC was studied by comparing the cyclic voltammogram recorded before and after cycling the potential between 0 and 0.6 V at 50 mV s<sup>-1</sup> for 10000 cycles. The EDLC was controlled and data recorded by an Autolab PGSTAT100 potentiostat (Metrohm).

A flexible and transparent EDLC prototype from SWCNT films was assembled as follows: a SWCNT film on 3x3 cm polyethylene terephthalate (PET) substrate was used as an electrode. A gel electrolyte was prepared by mixing a 5 wt% polyvinyl alcohol (PVA) solution (1 g in 10 mL de-ionized water) and concentrated H<sub>3</sub>PO<sub>4</sub> (2 g of 85% H<sub>3</sub>PO<sub>4</sub>) for 1 h. The viscous PVA/H<sub>3</sub>PO<sub>4</sub> solution was then applied on the SWCNT film with a brush to introduce a thin layer of PVA/ H<sub>3</sub>PO<sub>4</sub> on the SWCNT film surface. Two of these SWCNT films were then assembled face to face. Excess water in the PVA/H<sub>3</sub>PO<sub>4</sub> gel on the SWCNT surface was eliminated by leaving the supercapacitor in a fume hood overnight. After water had been evaporated, the PVA/ H<sub>3</sub>PO<sub>4</sub> was solidified and bound the electrodes together.

### 3. Results and discussion

#### 3.1 Properties of SWCNT and their thin films

In Fig. 1, TEM images of SWCNT are presented. Iron deposits from ferrocene used in the synthesis are visible and the iron content of the SWCNTs is 27 mol% from EDX. The SWCNTs have high specific surface area ( $950 \text{ m}^2 \text{ g}^{-1}$ ) making them suitable for EDLC applications.



**Figure 1.** A TEM image of the SWCNTs.

In Fig. 2a the Raman spectrum of the SWCNTs is presented. The D-band around  $1350 \text{ cm}^{-1}$  corresponds to defective structures while the G-band around  $1580 \text{ cm}^{-1}$  is indicative of graphitic structures [28]. The low intensity ratio between D and G bands ( $I_D/I_G = 0.029$ ) shows that structure of the SWCNT is highly ordered and of good quality.

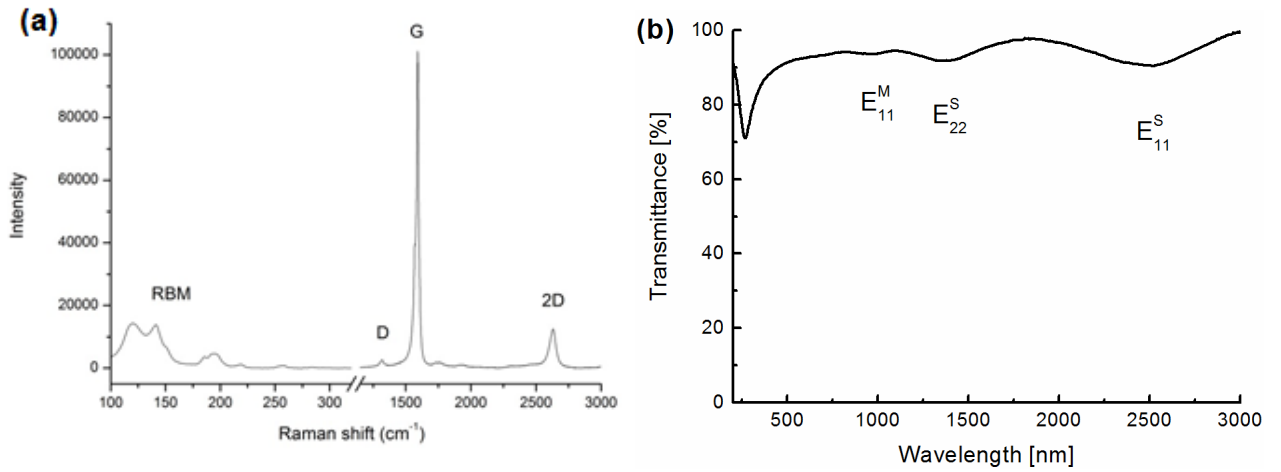
The optical transmittance of the SWCNT film (Fig. 2b) is extremely high (77-99 %) in the visible light and near infra-red regions (250-2250 nm). Hereinafter, we refer to the transmittance measured in the middle of visible range at 550 nm, which is 92% for this film. In the transmission spectrum electronic transitions in semiconducting (S11 and S22) and metallic (M11) SWCNTs, so called Van



Hove singularities, can be clearly observed. The mean diameter of SWCNTs estimated from the spectrum is about 2.2 nm [29], which corresponds to the TEM image. The roughness of the film is 30 nm [30] and the mass density of the film was determined to be  $3.10 \mu\text{g cm}^{-2}$ . The mass density of the other films (Table 2) was estimated from their transparency:

$$m_{x\%} = m_{92\%} \frac{\log(T_{x\%})}{\log(T_{92\%})} \quad (1)$$

where  $m_{x\%}$  is the film mass ( $\mu\text{g cm}^{-2}$ ) with transparency  $T_{x\%}$ .



**Figure 2.** Properties of the SWCNT films. (a) Raman spectrum of the SWCNT film with 633 nm excitation wavelength laser. (b) Optical transmittance (92% at 550 nm) of the SWCNT film with mass density  $3.10 \mu\text{g cm}^{-2}$ .

### 3.2 Electrochemical performance of the EDLC based on the SWCNT thin films

The electrochemical performance of the EDLC made from the SWCNT thin films with 92% transparency was studied in a two-electrode cell (Fig. 3 a-d). The cyclic voltammograms (Fig. 3a) show a rectangular shape that is ideal for EDLCs. The maximum potential ( $V_{\text{max}}$ ) is relatively low,

0.6 V, due to Faradaic processes starting at higher potentials. This is likely due to the Fe catalyst (27 mol%) remaining in the SWCNT film which can catalyze H<sub>2</sub> evolution from water [31]. The specific capacitance of one SWCNT thin film electrode was determined from constant current charging (Fig. 3b) at 0.53 A g<sup>-1</sup> between 0.5V<sub>max</sub> and V<sub>max</sub> [32] from the formulas:

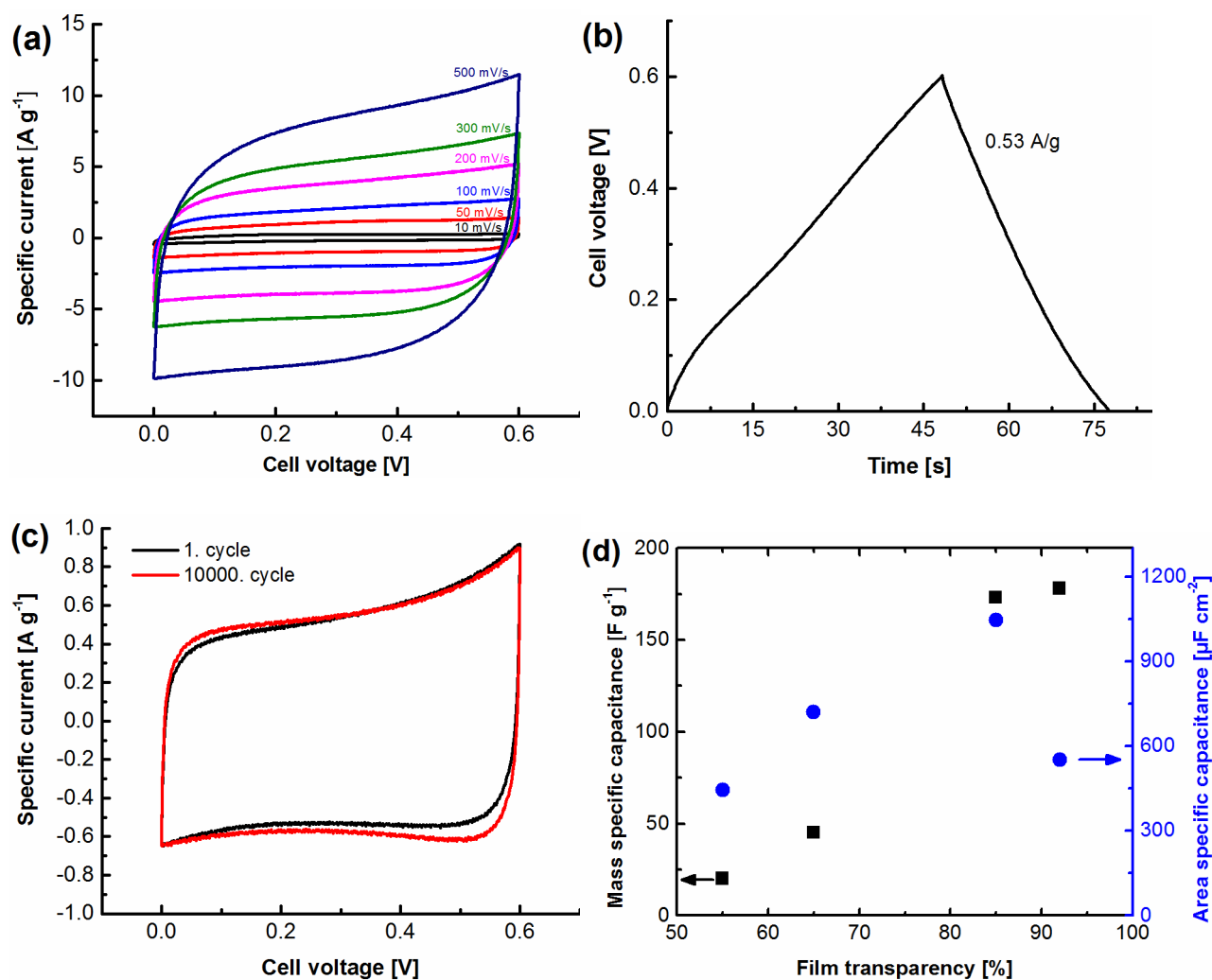
$$C_{\text{sp, mass}} = 4 \frac{I}{(\Delta V / \Delta t) m_t} \quad (2)$$

$$C_{\text{sp, area}} = 2 \frac{I}{(\Delta V / \Delta t) A_{\text{geom}}} \quad (3)$$

where  $I$  is absolute current,  $\Delta V$  the change in voltage,  $\Delta t$  the change in time,  $m_t$  is the total mass of two electrodes and  $A_{\text{geom}}$  is the geometric area of the supercapacitor (1 cm<sup>2</sup>). The multiplying factors 4 and 2 are derived from normalizing the specific capacitance for one electrode.

This yielded specific capacitances of 178 F g<sup>-1</sup> or 552 μF cm<sup>-2</sup> for a single SWCNT electrode. When the specific capacitance is calculated only relative to the carbon mass of the SWCNT, the mass specific capacitance is naturally even higher (482 F g<sup>-1</sup>, Table 2). The specific capacitance per electrode mass is extremely high and better than in other recent transparent and flexible EDLCs based on carbon nanomaterials presented in Table 1. Remarkably, considering the low amount of carbon material used in the electrodes (3.10 μg cm<sup>-2</sup>), the area specific capacitance is comparable or larger than with other EDLCs presented in Table 1. This is likely due to the high surface area and graphitization of the SWCNTs (Fig. 2a) [33]. In Fig. 3c, the cyclic voltammograms of the EDLC before and after 10000 cycles between 0 and 0.6 V are shown and surprisingly, a small improvement in current is seen. This is most likely due to better wetting of the SWCNT films over time and indicates excellent stability for the SWCNT electrodes compared to other carbon EDLCs [11,14,15]. In combination with the very high optical transparency of SWCNT films, these properties make our SWCNT films excellent material for flexible and transparent EDLC electrodes.

It is known that very thin films show high capacitances due to facile electrolyte transport [34-37]. Thus, when the film thickness is increased, the specific capacitance decreases to  $20 \text{ F g}^{-1}$  at 55% transparency corresponding to  $22.2 \text{ } \mu\text{g cm}^{-2}$  (Fig. 3d) showing the effect of longer electrolyte diffusion and electron transport through the layer. However, the specific capacitance normalized with the geometric area goes through a maximum at 85% transparency due to the mass specific capacitance decreasing while the mass of the electrode is increasing indicating that by sacrificing a small level of transparency, the area specific capacitance can be doubled and the SWCNTs used in applications demanding more energy. This enhancement is most likely caused by a smaller coverage of SWCNT at high transparency. Capacitance values of EDLCs containing different film thicknesses are presented in Table 2. The values are derived from constant current charging curves.



**Figure 3.** Electrochemical characterization of the SWCNT thin films as supercapacitors in 1 M  $\text{H}_2\text{SO}_4$ . (a) Cyclic voltammograms at various scan speeds (10, 50, 100, 200, 300 and 500  $\text{mV/s}$ ). (b) Constant current charging and discharging at  $0.53 \text{ A g}^{-1}$ . (c) Cyclic voltammogram of the supercapacitor before and after 10000 cycles between 0 and 0.6 V. (d) The effect of the SWCNT film thickness on the mass and area specific capacitance of the EDLC.

**Table 2.** The properties of different SWCNT films and the specific capacitances of the EDLCs made from them.

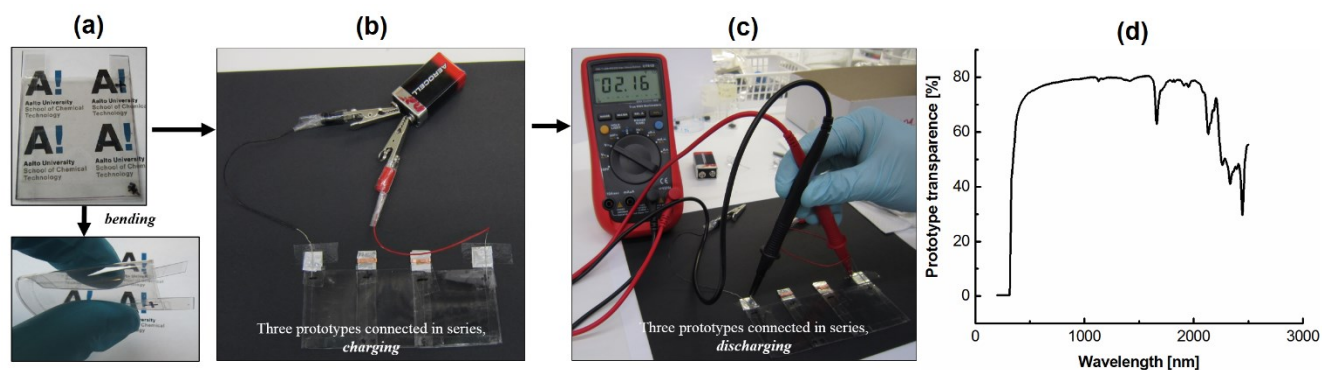
Transparency @ 550 nm (%)	Film mass $\mu\text{g cm}^{-2}$	$C_{\text{sp}}$ (total) <sup>a</sup> F g <sup>-1</sup>	$C_{\text{sp}}$ (carbon) <sup>b</sup> F g <sup>-1</sup>	$C_{\text{sp}}$ $\mu\text{F cm}^{-2}$
92	3.1	178	482	552
85	6.0	173	467	1045
65	16.0	45	121	720
55	22.2	20	54	444

<sup>a</sup> Specific capacitance of the SWCNT based on the total mass including iron residues

<sup>b</sup> Specific capacitance of the SWCNT based on the carbon mass

### 3.3 Transparent and flexible prototype EDLC

Finally, a flexible and transparent prototype was assembled from the SWCNT film with 92% transparency as electrodes, PET films as a casing and gel electrolyte. Its transparency and flexibility are demonstrated in Fig 4a. When three units were connected in series, total voltage of roughly 2.1 V could be achieved after charging for 1 minute (Fig. 4b-c). The total optical transparence spectrum of one prototype cell is presented in Fig. 4d and its value in the visible and near infra-red region is very high 75%. This prototype proves the feasibility of using the SWCNT films in flexible and transparent EDLCs. A more extensive study of the properties of these films under repeated bending is under way and the results will presented in a future paper.



**Figure 4.** (a) Photos of a transparent and flexible EDLC prototype based on the thin SWCNT films. (b) Charging and (c) discharging of three EDLC prototypes connected in series. (d) The optical transparency of single prototype cell in visible and near-infrared light.

#### 4. Conclusions

We have used a facile, one-step aerosol synthesis method to fabricate thin SWCNT films and used a simple dry press transfer technique to apply them to different substrates. The transparency of the films is 92% at 550 nm and they are highly transparent in the visible light and near infra-red wavelengths. The versatile transfer method allows easy combination of the SWCNT and substrate materials for transparent and flexible EDLCs. The films show extremely high mass specific capacitance of  $178 \text{ F g}^{-1}$  ( $482 \text{ F g}^{-1}$  calculated per mass of carbon) and area specific capacitance of  $552 \mu\text{F cm}^{-2}$  compared to other carbon based flexible and transparent EDLCs presented in literature (Table 1). The films are stable and their capacitance does not degrade during 10000 loading cycles. Finally, a transparent, flexible and stackable EDLC prototype was fabricated from the SWCNT films. All in all, the notable conductivity, transparency, capacitance and stability of these easily transferable SWCNT films hold great promise in transparent and flexible electronics applications.

## Acknowledgements

P.K. and T.K. thank Jane and Aatos Erkkö foundation for funding of this research. N.D.L and L.H.S thank MIDE at Aalto University for funding. A. G. N. and A. T. acknowledge the support by Ministry of Education and Science of the Russian Federation through project RFMEFI 58114X0006. This work made use of the facilities of Nanomicroscopy Center at Aalto University (Aalto-NMC).

## References

- [1] Simon P and Gogotsi Y 2008 *Nat Mater* **7** 845-54.
- [2] Gao K, Shao Z, Wu X, Wang X, Zhang Y, Wang W and Wang F 2013 *Nanoscale* **5** 5307-11.
- [3] Chen T, Peng H, Durstock M and Dai L 2014 *Sci Rep* **4** 3612.
- [4] Nakayama M, Okamura K, Athouël L, Crosnier O and Brousse T 2012 *ECS Trans* **41** 53-64.
- [5] Chen T, Xue Y, Roy AK and Dai L 2014 *ACS Nano* **8** 1039-46.
- [6] Fan X, Chen T and Dai L 2014 *RSC Adv* **4** 36996-7002.
- [7] Gao Y, Zhou YS, Xiong W, Jiang LJ, Mahjouri-samani M, Thirugnanam P, Huang X, Wang MM, Jiang L and Lu YF 2013 *APL Mater* **1** 012101.
- [8] Xu P, Kang J, Choi J, Suhr J, Yu J, Li F, Byun J, Kim B and Chou T 2014 *ACS Nano* **8** 9437-45.
- [9] Yu A, Roes I, Davies A and Chen Z 2010 *Appl Phys Lett* **96** 253105.
- [10] Chen P, Shen G, Sukcharoenchoke S and Zhou C 2009 *Appl Phys Lett* **94** 043113.
- [11] Lin H, Li L, Ren J, Cai Z, Qiu L, Yang Z and Peng H 2013 *Sci Rep* **3** 1353.
- [12] Jung HY, Karimi MB, Hahm MG, Ajayan PM and Jung YJ 2012 *Sci Rep* **2** 773.
- [13] Ge J, Cheng G and Chen L 2011 *Nanoscale* **3** 3084-8.
- [14] Niu Z, Zhou W, Chen J, Feng G, Li H, Hu Y, Ma W, Dong H, Li J and Xie S 2013 *Small* **9** 518-24.
- [15] Yuksel R, Sarioba Z, Cirpan A, Hiralal P and Unalan HE 2014 *ACS Appl Mater Interfaces* **6** 15434-9.

- [16] Yu M, Files BS, Arepalli S and Ruoff RS 2000 *Phys Rev Lett* **84** 5552-5.
- [17] Lau AK and Hui D 2002 *Composites Part B* **33** 263-77.
- [18] Sun D, Timmermans MY, Tian Y, Nasibulin AG, Kauppinen EI, Kishimoto S, Mizutani T and Ohno Y 2011 *Nat Nanotechnol* **6** 156-61.
- [19] Kaskela A, et al 2010 *Nano Lett* **10** 4349-55.
- [20] Nasibulin AG, Koltsova T, Nasibulina LI, Anoshkin IV, Semench A, Tolochko OV and Kauppinen EI 2013 *Acta Mater* **61** 1862-71.
- [21] Nasibulin AG, Moisala A, Brown DP, Jiang H and Kauppinen EI 2005 *Chem Phys Lett* **402** 227-32.
- [22] Nasibulin AG, Shandakov SD. Aerosol synthesis of carbon nanotubes. In: Agranovski I, editor. *Aerosols: Science and Technology*: Wiley-VCH; 2010. p. 65-89.
- [23] Tian Y, Timmermans MY, Partanen M, Nasibulin AG, Jiang H, Zhu Z and Kauppinen EI 2011 *Carbon* **49** 4636-43.
- [24] Moisala A, Nasibulin AG, Brown DP, Jiang H, Khriachtchev L and Kauppinen EI 2006 *Chem Eng Sci* **61** 4393-402.
- [25] Tian Y, Timmermans M, Kivistö S, Nasibulin A, Zhu Z, Jiang H, Okhotnikov O and Kauppinen E 2011 *Nano Res* **4** 807-15.
- [26] Nasibulin AG, Brown DP, Queipo P, Gonzalez D, Jiang H and Kauppinen EI 2006 *Chem Phys Lett* **417** 179-84.
- [27] Nasibulin AG, et al 2011 *ACS Nano* **5** 3214-21.
- [28] Dresselhaus MS, Dresselhaus G, Saito R and Jorio A 2005 *Phys Rep* **409** 47-99.
- [29] Gorkina AL, et al 2016 *Carbon* **100** 501-7.
- [30] Funde AM, et al 2016 *Nanotech* **27** 185401.
- [31] Du P and Eisenberg R 2012 *Energy Environ Sci* **5** 6012-21.
- [32] Stoller MD and Ruoff RS 2010 *Energy Environ Sci* **3** 1294-301.
- [33] Weingarth D, Zeiger M, Jäckel N, Aslan M, Feng G and Presser V 2014 *Adv Energy Mat* **4** 1400316.
- [34] Chmiola J, Largeot C, Taberna P, Simon P and Gogotsi Y 2010 *Science* **328** 480-3.
- [35] Hu L, Choi JW, Yang Y, Jeong S, La Mantia F, Cui L and Cui Y 2009 *PNAS* **106** 21490-4.
- [36] Hu S, Zhang S, Pan N and Hsieh Y 2014 *J Power Sources* **270** 106-12.



[37] Luo J, Jang HD and Huang J 2013 *ACS Nano* **7** 1464-71.